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and red, represent guesstimatepoints) Question 1 (a) (i)  $Pb^{2+} (aq)$

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NEXT PAGE. -7-  $CH_3CH_2COOH(aq) + H_2O(l) \rightleftharpoons CH_3CH_2COO^-(aq) + H_3O^+(aq)$

2. Propanoic acid,  $CH_3CH_2COOH$ , is a carboxylic acid that reacts with water according to the equation above. At

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(d) Calculate the number of moles of precipitate that is produced in the experiment.  $1.698 \text{ g} - 1.462 \text{ g} = 0.236 \text{ g PbI}_2 \text{ (s)}$   $\frac{0.236 \text{ g}}{461 \text{ g/mol}} = 5.12 \times 10^{-4} \text{ mol PbI}_2$

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at the beginning of the reaction and at the end, do not change in any way (i.e., they are spectators), and as such can be omitted from the equation.

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